

Ontario College of Pharmacy.

CHEMISTRY.

JUNE 17th, 1884.

Examiner—W. E. SAUNDERS.

1. Define the terms (a) Atomic Weight, (b) Efflorescence, (c) Monad, (d) Specific Gravity, (e) Destructive Distillation.
2. Give the specific gravity of (a) Bromine, (b) Liquor Hydrarg. Nit. Acid., (c) Glycerine, (d) Acid. Hydrochlor.
2.966, 2.246, 1.25, 1.18
3. Give tests for the purity of (a) Potassium Iodide, (b) Ferrous Sulphate, (c) Magnesium Sulphate, (d) Mercurous Iodide.
Ca₂ Stades, H₂O, ex + H₂S, Ca + metals, other for mercurio
4. Give the chemical formula of (a) Chloroform, (b) Stychnia, (c) Sulphate of Copper, (d) Purified Animal Charcoal, (e, f and g) Red, Yellow and Black Oxide of Mercury, and (h) explain the differences between the last three.
CHCl₃, C₁₇H₁₅N₃O₂, S₂CuO, H
5. Give a brief description of the Manufacture of Sulphate of Quinine.
6. Describe the manufacture of (a) Sulphuric and (b) Citric Acid.
7. Explain by equations the reactions which take place in the manufacture of (a) Liq. Ferri Pernit, and (b) Ferri Carb. Sacch.
- 8-10. Oral and specimens.